



CANDIDATE
NAME

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CENTRE
NUMBER

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CANDIDATE
NUMBER

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0620/32

February/March 2020

1 hour 15 minutes

You must answer on the question paper.

No additional materials are needed.

- Answer **all** questions.
- Use a black or dark blue pen. You may use an HB pencil for any diagrams or graphs.
- Write your name, centre number and candidate number in the boxes at the top of the page.
- Write your answer to each question in the space provided.
- Do **not** use an erasable pen or correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.
- You should show all your working and use appropriate units.

- The total mark for this paper is 80.
- The number of marks for each question or part question is shown in brackets [].
- The Periodic Table is printed in the question paper.

This document has **20** pages. Blank pages are indicated.

1 (a) A list of compounds is shown.

aluminium oxide
calcium bromide
calcium oxide
ethane
ethene
hydrogen chloride
methane
nitrogen dioxide
potassium iodide
potassium manganate(VII)
sodium chloride

Answer the following questions using only the compounds in the list.
Each compound may be used once, more than once or not at all.

Which compound:

(i) when in acidified solution, is used to test for sulfur dioxide

..... [1]

(ii) is the main constituent of natural gas

..... [1]

(iii) when dissolved in water, gives a yellow precipitate on addition of acidified aqueous silver nitrate

..... [1]

(iv) is used in flue gas desulfurisation to neutralise acidic gases

..... [1]

(v) is a reactant used in the manufacture of ethanol?

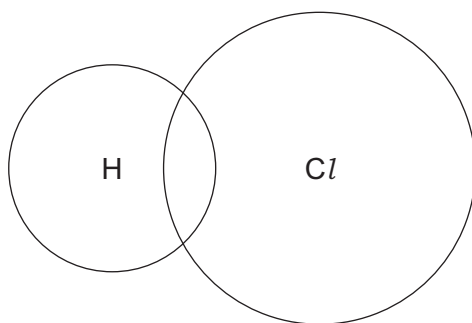
..... [1]

(b) What is the meaning of the term *compound*?

.....
..... [2]

(c) Complete the electronic structure of a molecule of hydrogen chloride.

Show only the outer shell electrons.

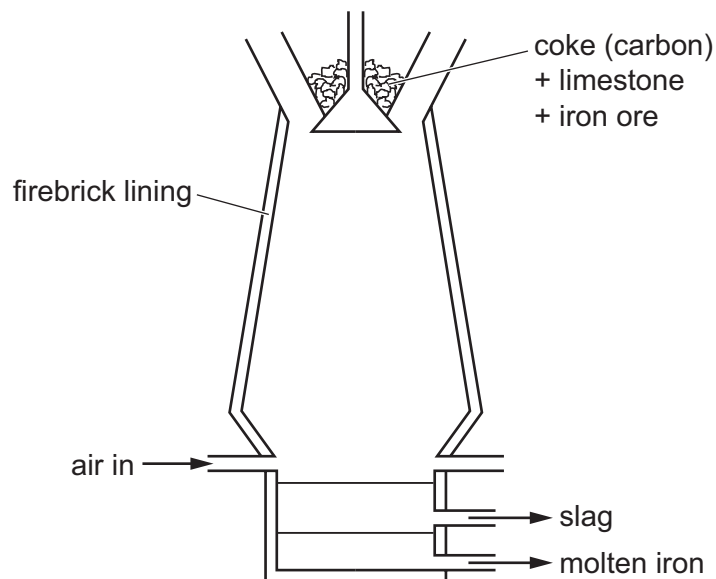


[2]

[Total: 9]

- 2 Iron is extracted by heating a mixture of coke (carbon), limestone and iron ore in air in a blast furnace.

A diagram of the blast furnace is shown.



- (a) Name the ore of iron added to the blast furnace.

..... [1]

- (b) The impurities in the iron ore are removed as slag.

- (i) What information in the diagram shows that slag is less dense than molten iron?

..... [1]

- (ii) Which **one** of the substances added to the blast furnace helps to remove the impurities?

Explain how it does this.

substance

explanation

.....

.....

[3]

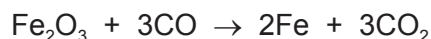
- (c) Hot air is blown into the blast furnace.

Explain why.

.....

..... [1]

(d) The chemical equation for one of the reactions in the blast furnace is shown.



(i) How does this equation show that Fe_2O_3 has been reduced?

..... [1]

(ii) When 16.0 g of Fe_2O_3 react with excess carbon monoxide, 11.2 g of iron are produced.

Calculate the mass of iron produced when 4.0 g of Fe_2O_3 react with excess carbon monoxide.

mass of iron = g [1]

(e) An isotope of iron is shown.



Deduce the number of electrons, protons and neutrons in an atom of this isotope of iron.

number of electrons

number of protons

number of neutrons

[3]

(f) Iron is a transition element.

Which **two** of these statements about iron are correct?

Tick **two** boxes.

Iron forms coloured compounds.

☐

Iron can act as a catalyst.

☐

Iron is brown when freshly cut.

☐

Iron has a low density.

☐

Iron has a low melting point.

☐

[2]

[Total: 13]

3 Water is essential for many industrial processes.

(a) State **one** use of water in industry.

..... [1]

(b) What is the pH of pure water?

Draw a circle around the correct answer.

pH 0

pH 6

pH 7

pH 14

[1]

(c) Filtration and chlorination are two of the steps used in water treatment.

Describe the purpose of each of these steps.

filtration

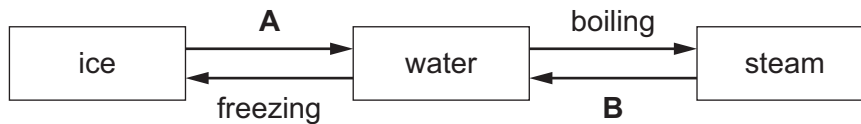
.....

chlorination

.....

[2]

(d) The changes of state of water are shown.



Give the names of the changes of state represented by **A** and **B**.

A

B

[2]

(e) The table compares the reactions of four metals with both steam and dilute hydrochloric acid.

metal	reaction with steam at 200 °C	observation with dilute hydrochloric acid
copper	no reaction	no bubbles formed
magnesium	rapid reaction	bubbles form rapidly
nickel	no reaction	bubbles form slowly
zinc	rapid reaction	bubbles form slowly

Put the four metals in order of their reactivity.
Put the least reactive metal first.

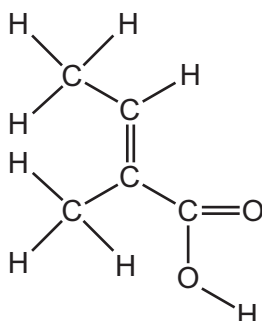
least reactive  most reactive

[2]

[Total: 8]

- 4 Angelic acid and ethanoic acid are both carboxylic acids.

The structure of angelic acid is shown.



- (a) (i) On the structure of angelic acid, draw a circle around the functional group that shows that it is a carboxylic acid. [1]

- (ii) Deduce the formula of angelic acid to show the number of carbon, hydrogen and oxygen atoms.

..... [1]

- (iii) Angelic acid is an unsaturated compound.

Describe a chemical test to distinguish between an unsaturated and a saturated compound.

test

result with unsaturated compound

result with saturated compound

[3]

- (b) The formula of ethanoic acid is $C_2H_4O_2$.

Complete the table to calculate the relative molecular mass of ethanoic acid.

Use the Periodic Table to help you.

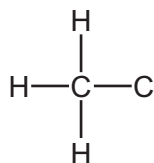
type of atom	number of atoms	relative atomic mass	
carbon	2	12	$2 \times 12 = 24$
hydrogen			
oxygen		16	

relative molecular mass =

[2]

- (c) Ethanoic acid can be reduced to ethanol.

Complete the structure of ethanol to show all of the atoms and all of the bonds.



[1]

- (d) Ethanol can be manufactured by fermentation.

Describe the process of fermentation to include:

- the names of the reactants and catalyst

.....
.....

- the conditions required

.....
.....
.....

- the name of the process used to separate the ethanol from the rest of the reaction mixture.

.....

[4]

[Total: 12]

- 5 The table shows some properties of four metals in Group I of the Periodic Table.

metal	melting point /°C	boiling point /°C	relative electrical conductivity
sodium	98	883	
potassium	63	760	14
rubidium		686	8
caesium	29	669	5

- (a) Complete the table to estimate:

- the melting point of rubidium
- the relative electrical conductivity of sodium.

[2]

- (b) What is the physical state of caesium at 20 °C?

Give a reason for your answer.

.....

 [2]

- (c) Describe the trend in the boiling points of the Group I metals.

..... [1]

- (d) When potassium reacts with water, a coloured flame is seen and a gas is produced that pops with a lighted splint.

- (i) Complete the chemical equation for this reaction.



- (ii) State the colour of the flame when potassium reacts with water.

..... [1]

- (iii) The solution formed is alkaline.

Describe how you can use universal indicator solution to determine the pH of the solution.

.....
 [2]

[Total: 10]

6 This question is about the reactions of magnesium with nitric acid.

(a) The equation for the reaction of magnesium with concentrated nitric acid is shown.



(i) The reaction is exothermic.

What is the meaning of the term *exothermic*?

..... [1]

(ii) Which word best describes the compound $\text{Mg}(\text{NO}_3)_2$?

Draw a circle around the correct answer.

acid base oxide salt [1]

(iii) Oxides of nitrogen are formed when fossil fuels are burned.

What type of chemical reaction occurs when fossil fuels are burned?

Draw a circle around the correct answer.

combustion cracking fermentation neutralisation [1]

(iv) Oxides of nitrogen dissolve in rain water to form acid rain.

State **one** adverse effect of acid rain on buildings.

..... [1]

(b) When very dilute nitric acid reacts with magnesium powder, hydrogen is produced.

(i) Describe a practical method for investigating the rate of this reaction.

.....

 [3]

(ii) What effect would each of the following have on the rate of this reaction?

- Larger pieces of magnesium are used instead of magnesium powder.

All other conditions stay the same.

.....

- The temperature of the reaction mixture is increased.

All other conditions stay the same.

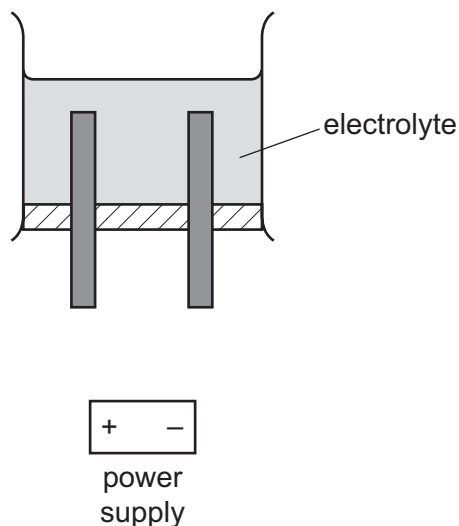
.....

[2]

[Total: 9]

7 (a) The electrolysis of dilute sulfuric acid produces gases at both electrodes.

(i) The incomplete apparatus is shown.



Complete the diagram by:

- labelling the anode and cathode
- adding connecting wires
- showing how the gases are collected.

[3]

(ii) Name the products formed at each electrode.

positive electrode

negative electrode

[2]

(b) Carbon dioxide is produced when sulfuric acid reacts with sodium carbonate.

Name the **two** other products which are formed.

..... and

[2]

(c) Describe the test for carbon dioxide.

test

observations

[2]

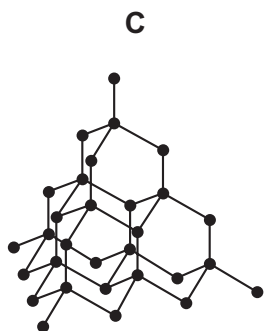
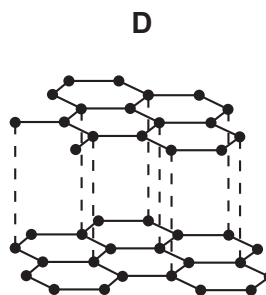
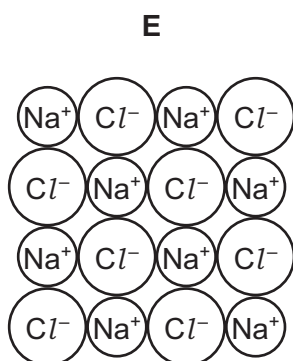
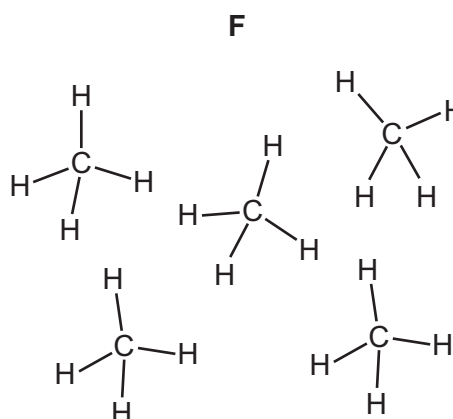
(d) Carbon dioxide is a greenhouse gas.

State **one** effect of greenhouse gases on the environment.

..... [1]

[Total: 10]

- 8 (a) The structures of four substances **C**, **D**, **E** and **F**, are shown.



- (i) Which **one** of these substances, **C**, **D**, **E** or **F**, is a gas at room temperature?

..... [1]

- (ii) What type of bonding is present in substance **E**?

..... [1]

- (iii) Which **one** of these substances, **C**, **D**, **E** or **F**, is soluble in water?

..... [1]

- (iv) Which **one** of these substances, **C**, **D**, **E** or **F**, conducts electricity when solid?

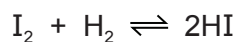
..... [1]

- (b) The halogens have molecules containing two atoms.

What is the name for molecules containing only two atoms?

..... [1]

(c) The reaction of iodine with hydrogen is shown.



What is the meaning of the symbol $\rightleftharpoons$?

..... [1]

(d) Iodine is formed when chlorine reacts with aqueous potassium iodide.

(i) Complete the chemical equation for this reaction.



(ii) When aqueous iodine is mixed with aqueous potassium chloride, there is no reaction.

Suggest, in terms of chemical reactivity, why there is no reaction.

..... [1]

[Total: 9]

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Group														
I	II							III	IV	V	VI	VII	VIII	
		<div>Key<div>atomic number atomic symbol name relative atomic mass</div></div>							<div>1Hhydrogen1</div>					2Hehelium4
3Li lithium 7	4Be beryllium 9								5B boron 11	6C carbon 12	7N nitrogen 14	8O oxygen 16	9F fluorine 19	
11Na sodium 23	12Mg magnesium 24								13Al aluminium 27	14Si silicon 28	15P phosphorus 31	16S sulfur 32	17Cl chlorine 35.5	
19K potassium 39	20Ca calcium 40								31Ga gallium 70	32Ge germanium 73	33As arsenic 75	34Se selenium 79	35Br bromine 80	
37Rb rubidium 85	38Sr strontium 88								49In indium 115	50Sn tin 119	51Sb antimony 122	52Te tellurium 128	53I iodine 127	
55Cs caesium 133	56Ba barium 137								81Tl thallium 204	82Pb lead 207	83Bi bismuth 209	84Po polonium —	85At astatine —	
87Fr francium —	88Ra radium —								112Cn copernicium —	114Fl flerovium —	116Lv livermorium —	118Og oganesson —	119Ts tennessine —	
103Nh nihonium —	104Fl flerovium —								115Mc moscovium —	117Ts tennessine —	119Og oganesson —	120Uue unbinilium —	121Uuh ununilium —	

lanthanoids

actinoids

57	La	lanthanum	139	58	Ce	cerium	140	59	Pr	praseodymium	141	60	Nd	neodymium	144	61	Pm	promethium	—	62	Sm	samarium	150	63	Eu	euporium	152	64	Gd	gadolinium	157	65	Tb	terbium	159	66	Dy	dysprosium	163	67	Ho	holmium	165	68	Er	erbium	167	69	Tm	thulium	169	70	Yb	ytterbium	173	71	Lu	lutetium	175
89	Ac	actinium	—	90	Th	thorium	232	91	Pa	protactinium	231	92	U	uranium	238	93	Np	neptunium	—	94	Pu	plutonium	—	95	Am	americium	—	96	Cm	curium	—	97	Bk	berkelium	—	98	Cf	californium	—	99	Es	einsteinium	—	100	Fm	femium	—	101	Md	merendelevium	—	102	No	nobelium	—	103	Lr	lawrencium	—